

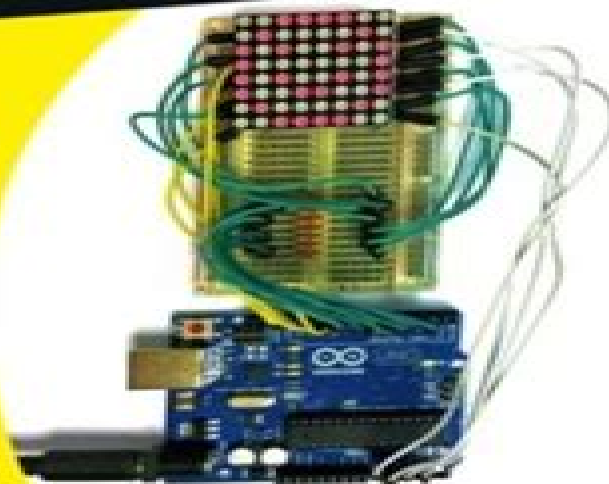
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web solution if we compare the time that has passed to the isotope's half life we note that 44.0 s is exactly 4 half lives so using the previous expression $n = 4$ substituting and solving results in the following less than one third of a gram of fluorine-20 remains test yourself the half life of titanium-44 is 60.0 y

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web learning objectives by the end of this section you will be able to define half life define dating calculate age of old objects by radioactive dating unstable nuclei decay however some nuclides decay faster than others for example radium and polonium discovered by the Curies decay faster than uranium

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for cobalt 60 which has a half life of 5 27 years 50 remains after 5 27 years one half life 25 remains after 10 54 years two half lives 12 5 remains after 15 81 years three half lives and so on

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web we utilize the equation that relate amount remaining initial mass and number of half lives $n = \frac{\ln \left(\frac{N_0}{N} \right)}{\ln 2}$ $N = N_0 \left(\frac{1}{2} \right)^n$ 125 g 5 what is the half life of an isotope that is 80 remained after 16 days solution remaining 80 100 therefore $n = \frac{\ln \left(\frac{100}{80} \right)}{\ln 2}$ now using the half life equation in example 1 we have

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