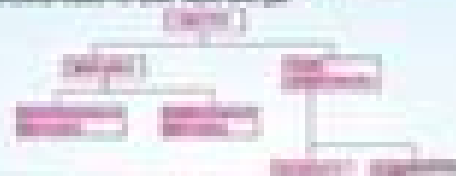


1. CLASSIFICATION OF MATTER

- Matter can exist in three physical states: solid, liquid and gas.



- Various sub-classifications of matter are given below:
- Matter is classified into pure substances and mixtures.
- Pure substances are further classified into elements and compounds.
- Compounds are further classified into ionic compounds and covalent compounds.
- Compounds are further classified into organic and inorganic compounds.
- Compounds are further classified into simple and complex compounds.

2. PROPERTIES OF MATTER AND THEIR MEASUREMENT

- Properties of matter are classified into physical and chemical properties.
- Physical properties are those which can be observed or measured without changing the identity of the substance.
- Chemical properties are those which can be observed or measured by changing the identity of the substance.
- Colour, odour, melting point, boiling point, etc. are physical properties.
- Combustion, corrosion, etc. are chemical properties.
- Quantitative measurement of physical properties is called measurement of matter.
- Measurement of matter is called measurement.

- Some important quantities and their units:

Quantity	SI Unit
Mass (g)	Kilogram (kg)
Volume (cm ³)	Cubic metre (m ³)
Temperature (°C)	Kelvin (K)

Some important units and their symbols

10 ³ times	10 ⁶ times	10 ⁹ times	10 ¹² times	10 ¹⁵ times
kilo (k)	mega (M)	giga (G)	tera (T)	peta (P)

Some important units

- 1 kg = 1000 g, 1 g = 1000 mg, 1 mg = 1000 µg, 1 µg = 1000 ng, 1 ng = 1000 pg, 1 pg = 1000 fg.

3. UNCERTAINTY IN MEASUREMENT

- Measurement is a comparison of a quantity with a standard unit.
- Measurement is always uncertain.
- Measurement is always uncertain.

- Significant figures are those digits which are known with certainty plus one additional uncertain digit.

- Rules for determining the number of significant figures:

- Non-zero digits are significant.
- Zeros between two non-zero digits are significant.
- Zeros at the end or right of a number are significant if they are on the right side of a decimal point.
- Leading zeros are not significant.
- In a number written in scientific notation, all digits are significant.
- Exact values are the values of measurements for which there is no uncertainty.
- Round-off is the agreement of a measured value to the true value of the result.
- In addition and subtraction of significant figures, the result should have the same number of significant figures as the least number of significant figures in the original numbers.
- In multiplication and division of significant figures, the result should have the same number of significant figures as the least number of significant figures in the original numbers.

Aakash Chemistry Lab Manual

Léo Pomar



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